



Standard Test Method for Conducting Cyclic Galvanostaircase Polarization¹

This standard is issued under the fixed designation G 100; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This test method gives a procedure for conducting cyclic galvanostaircase polarization (GSCP) to determine relative susceptibility to localized corrosion (pitting and crevice corrosion) for aluminum alloy 3003-H14 (UNS A93003) (1).² It may serve as guide for examination of other alloys (2–5). This test method also describes a procedure that can be used as a check for one's experimental technique and instrumentation.

1.2 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 ASTM Standards:

D 1193 Specification for Reagent Water³

G 1 Practice for Preparing, Cleaning, and Evaluating Corrosion Test Specimens⁴

G 5 Reference Test Method for Making Potentiostatic and Potentiodynamic Anodic Polarization Measurements⁴

G 59 Practice for Conducting Potentiodynamic Polarization Resistance Measurements⁴

G 69 Practice for Measurement of Corrosion Potentials of Aluminum Alloys⁴

3. Significance and Use

3.1 In this test method, susceptibility to localized corrosion of aluminum is indicated by a protection potential (E_{prot}) determined by cyclic galvanostaircase polarization (1). The more noble this potential, the less susceptible is the alloy to initiation of localized corrosion. The results of this test method are not intended to correlate in a quantitative manner with the rate of propagation of localized corrosion that one might observe in service.

3.2 The breakdown (E_b), and protection potentials (E_{prot}) determined by the cyclic GSCP method correlate with the constant potential corrosion test (immersion-glassware) result

for aluminum (1, 6, 8). When the applied potential was more negative than the GSCP E_{prot} , no pit initiation was observed. When the applied potential was more positive than the GSCP E_{prot} , pitting occurred even when the applied potential was less negative than E_b .

3.2.1 Severe crevice corrosion occurred when the separation of E_b and E_{prot} was 500 mV or greater and E_{prot} was less than -400 mV Vs. SCE (in 100 ppm NaCl) (1, 6, 7). For aluminum, E_{prot} determined by cyclic GSCP agrees with the repassivation potential determined by the scratch potentiostatic method (1, 10). Both the scratch potentiostatic method and the constant potential technique for determination of E_{prot} require much longer test times and are more involved techniques than the GSCP method.

3.3 DeBerry and Viebeck (3–5) found that the breakdown potentials (E_b) (galvanodynamic polarization, similar to GSCP but no kinetic information) had a good correlation with the inhibition of localized corrosion of 304L stainless steel by surface active compounds. They attained accuracy and precision by avoiding the strong induction effect which they observed by the potentiodynamic technique.

3.4 If this test method is followed using the specific alloy discussed it will provide (GSCP) measurements that will reproduce data developed at other times in other laboratories.

3.5 E_b and E_{prot} obtained are based on the results from eight different laboratories that followed the standard procedure using aluminum alloy 3003-H14 (UNS A93003). E_b and E_{prot} are included with statistical analysis to indicate the acceptable range.

4. Apparatus

4.1 *Cell*—The cell should be constructed of inert materials such as borosilicate glass and PTFE fluorocarbon. It should have ports for the insertion of a working electrode (1 cm² flat specimen holder (Note 1) is very convenient), two auxiliary electrodes, salt bridge for reference electrode, and a thermometer or a thermostat probe for temperature control. The figure in Test Method G 5 would be satisfactory, but a flat bottom cell is also satisfactory provided that all of the essential ports are provided. (See Ref (9) for details.)

NOTE 1—These specific recommendations and conditions were followed to improve the inter-laboratory precision during the round robin for galvanostaircase polarization.

4.2 *Current Staircase Generator and Recorder*—The schematic diagram of the apparatus is given in Fig. 1. The recorder

¹ This test method is under the jurisdiction of ASTM Committee G-1 on Corrosion of Metals and is the direct responsibility of Subcommittee G1.11 on Electrochemical Measurements in Corrosion Testing.

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² The boldface numbers in parentheses refer to the list of references at the end of this test method.

³ *Annual Book of ASTM Standards*, Vol 11.01.

⁴ *Annual Book of ASTM Standards*, Vol 03.02.

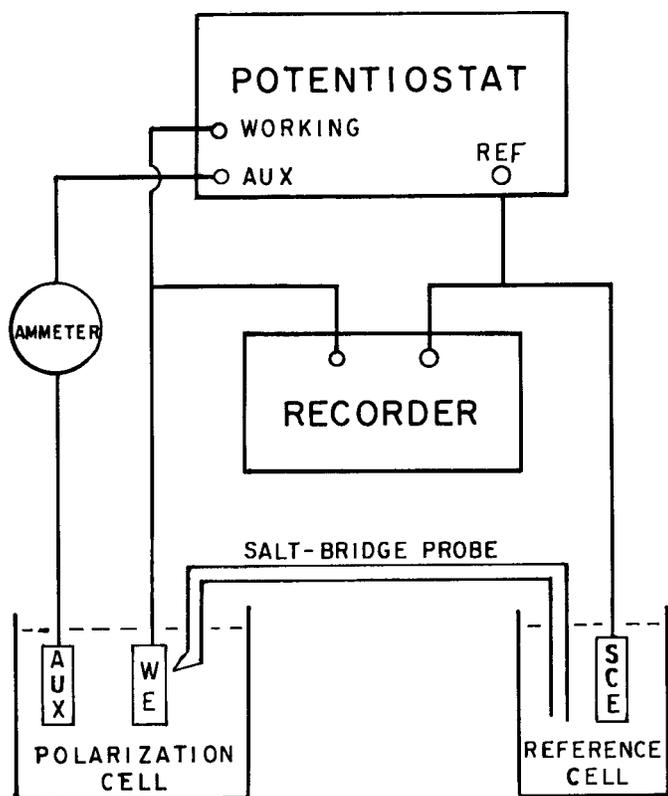


FIG. 1 Schematic Wiring Diagram for Galvanostaircase Polarization

may be replaced by a plotter if the current staircase signal is generated with the aid of a computer. The current staircase may be generated manually (Note 2) but this is not recommended. The most convenient current staircase generators are found in recent commercial potentiostats where the software is available. The electrical equipment may be checked in accordance with the procedure in Practice G 59.

NOTE 2—The current staircase signal was generated manually in the round robin because automated system or software was not available when this project was started.

4.3 Electrodes:

4.3.1 Working Electrode—For generating data to be compared to the reference data included herein, use type 3003-H14 (UNS A93003) A1 in sheet form. Cut 1.55 cm diameter circles and prepare in accordance with Practice G 1 using 600-grit diamond slurry on a flat lapping machine. Install in flat specimen holder using PTFE gasket (no crevice type) (Note 1) so that 1 cm² is exposed to the test solution. Apply 28.8 m-g (40 in. - oz) of torque.

4.3.2 Auxiliary Electrodes—Graphite, (ultrafine grade) (Note 3).

NOTE 3—Coarse grades of graphite should be avoided because they absorb solute impurities. Ultrafine grades are available from spectrographic supply companies.

4.3.3 Reference Electrode—Saturated calomel (Note 1). It should be checked against another reference which has not been exposed to test solutions and they should be within 3 mV of each other. Practice G 69 round robin test conducted by

G01.11 (unpublished results) indicate that potential difference should not exceed 2 or 3 mV. The reference electrode is connected to the test bridge solution which consists of 75 % saturated KCl, prepared by adding 1 part (by volume) of distilled water to 3 parts saturated KCl. When the bridge is in active use, the bridge solution should be replaced once each day and the bridge tip immersed in this solution when not in use. Any test solution that does not deposit films may also be used in the bridge. (The VYCOR⁵ tip **should not** be allowed to go to dryness.)

4.4 Magnetic Stirrer.

5. Procedure

5.1 Test solution, 3000 ± 30 ppm (0.0513 M) NaCl. For example, transfer 6.000 g reagent grade NaCl to a 2-L volumetric flask. Dissolve in ASTM Type IV water (demineralized or distilled) and dilute to the mark. (See Specification D 1193.)

5.2 Assemble cell with the electrodes described in Section 4. Place the reference bridge probe about 2 probe tip diameters away from the working electrode.

5.3 Fill the cell with the test solution so that the level is about 25 mm (1 in.) above the working electrode.

5.4 Maintain a temperature of 25 ± 1°C.

5.5 Do not deaerate.

5.6 Turn on the magnetic stirrer to a maximum speed that will maintain a smooth vortex above the specimen without whipping air bubbles into the solution.

5.7 Apply a current staircase signal from 0 to 120 μA/cm² using a step height of 20 μA/cm² and step duration of 2 min; reverse the current staircase scan to 0 current. Record the voltage transients on an X-Y or X-T recorder or plotter as shown in Fig. 2 (Note 4). In order to differentiate between the steady-state potential values of the forward scan from those of the reverse scan, it would be helpful to (1) delay the actual reversal of the pen about 12 s after dropping from 120 to 100 μA/cm² so that there will be a separation of about 24 s between the forward and reverse steady state points and (2) change the pen color in the reverse scan.

⁵ VYCOR is a trademark of Owen Corning, Code No. 7930 glass.

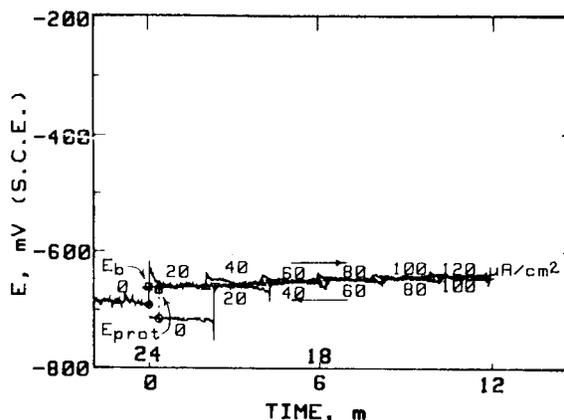


FIG. 2 Cyclic GSCP Curve of 3003 A1 in 3000 ppm NaCl (Taken from Ref. 7)

NOTE 4—Fig. 2 can be elucidated with the help of Fig. 3. The upper curve in Fig. 3 shows the current staircase signal applied in 5.7 and the lower curve gives schematic voltage response transients with the current density given for each transient. The current is selected at the end of each step (even though current is constant during a step) because the steady state voltage is obtained at the end of the step. This allows extrapolation to zero current which is a discrete current value at each end of the lower curve. In Fig. 2, the down-steps are reversed with a slight delay to separate up-step (triangles pointing upward) from the down-step (triangle pointing downward) steady state voltage.

5.8 *Extrapolation*—Extrapolate the up-step points to zero current to obtain E_b . Similarly, extrapolate the down-step points to obtain E_{prot} . Fig. 3 and Fig. 2 give examples of these extrapolations.

6. Precision and Bias ⁶

6.1 *Precision*—The precision information is based on data obtained by the GSCP Task Group with eight laboratories participating. Each laboratory ran duplicate results on the one test solution. The mean value for E_b was -636 mV with a standard deviation of 15.8 mV. The mean value for E_{prot} was -652 mV with a standard deviation of 14.8 .

6.2 The repeatability of this technique was 3.5 mV for E_{prot} and 7.3 mV for E_b in terms of the pooled standard deviation. (See Note 5.)

6.3 *Bias*—This procedure has no bias because the values of E_b and E_{prot} can be defined only in terms of this method. If the voltage transients are omitted from Fig. 3 and Fig. 2, typical quasi-stationary galvanostatic polarization plots are obtained. However, the kinetic and noise information derived from the voltage transients are desirable attributes of GSCP.

NOTE 5—The standard deviation was derived from

$$S^2 = \sum_{i=1}^N \frac{(Y_i - \bar{Y})^2}{N-1} \tag{1}$$

where:

Y = the i^{th} result,

\bar{Y} = the average of all Y_i values, and

N = is the total number of results.

The pooled standard deviation was derived from

$$(\text{Spooled})^2 = \sum_{i=1}^K \frac{(J_{1i} - J_{2i})^2}{2K} \tag{2}$$

where:

K = the number of laboratories and J_{1i} and J_{2i} are the duplicate results from the i^{th} laboratory.

7. Keywords

7.1 aluminum; corrosion; electrochemical measurement; galvanostaircase; localized corrosion; polarization

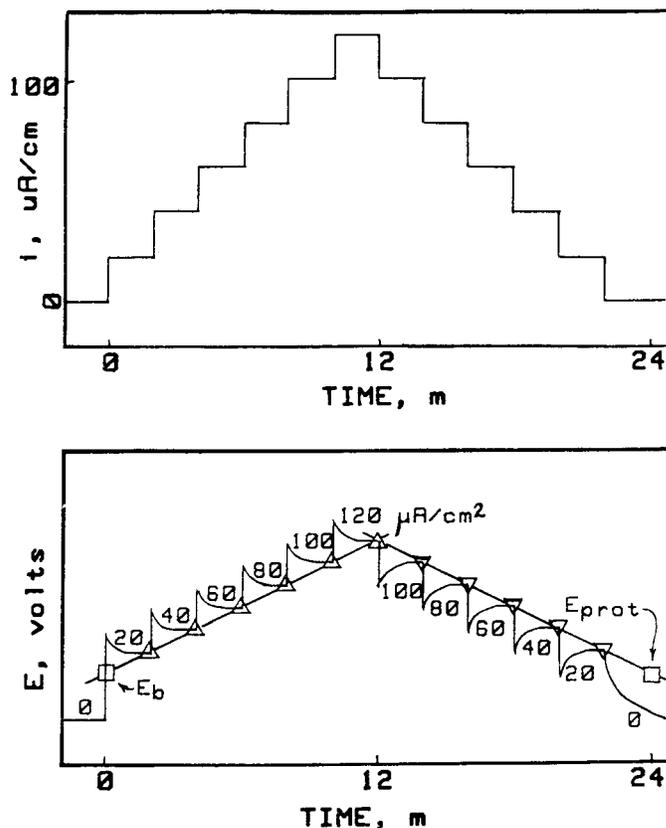


FIG. 3 Relationship of a Schematic GSCP Curve (lower) to the Current Staircase Signal (upper)

⁶ The results of the round robin are available from ASTM headquarters.

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